

Redox Reaction Practice Pre-AP

Complete on a separate sheet of paper using the steps provided to you in the tutorial. Write final answer below equation on this sheet.

- $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{H}^+ \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+} + \text{H}_2\text{O}$
- $\text{Cr}(\text{OH})_4^- + \text{OH}^- + \text{H}_2\text{O}_2 \rightarrow \text{CrO}_4^{2-} + \text{H}_2\text{O}$ (Basic)
- $\text{MnO}_2 + \text{H}^+ + \text{NO}_2^- \rightarrow \text{NO}_3^- + \text{Mn}^{2+} + \text{H}_2\text{O}$
- $\text{Sn}(\text{OH})_3^- + \text{Bi}(\text{OH})_3 + \text{OH}^- \rightarrow \text{Sn}(\text{OH})_6^{2-} + \text{Bi}$ (Basic)
- $\text{Al} + \text{NO}_3^- + \text{OH}^- + \text{H}_2\text{O} \rightarrow \text{Al}(\text{OH})_4^- + \text{NH}_3$ (Basic)
- $\text{NO}_2 + \text{OH}^- \rightarrow \text{NO}_3^- + \text{NO}_2^- + \text{H}_2\text{O}$ (Basic)
- $\text{MnO}_4^- + \text{H}_2\text{O} + \text{NO}_2^- \rightarrow \text{MnO}_2 + \text{NO}_3^- + \text{OH}^-$ (Basic)
- $\text{I}^- + \text{H}^+ + \text{NO}_2^- \rightarrow \text{NO} + \text{H}_2\text{O} + \text{I}_2$
- $\text{Hg}_2\text{Cl}_2 + \text{NH}_3 \rightarrow \text{Hg} + \text{HgNH}_2\text{Cl} + \text{NH}_4^+ + \text{Cl}^-$ (Basic)
- $\text{CrO}_4^{2-} + \text{OH}^- + \text{HSnO}_2 \rightarrow \text{CrO}_2^- + \text{H}_2\text{O} + \text{HSnO}_3^-$ (Basic)
- $\text{C}_2\text{H}_4 + \text{MnO}_4^- + \text{H}^+ \rightarrow \text{CO}_2 + \text{Mn}^{2+} + \text{H}_2\text{O}$
- $\text{H}_2\text{S} + \text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow \text{Cr}^{3+} + \text{S} + \text{H}_2\text{O}$
- $\text{ClO}_3^- + \text{H}_2\text{O} + \text{I}_2 \rightarrow \text{IO}_3^- + \text{Cl}^- + \text{H}^+$
- $\text{Cu} + \text{H}^+ + \text{SO}_4^{2-} \rightarrow \text{Cu}^{2+} + \text{H}_2\text{O} + \text{SO}_2$
- $\text{H}_2\text{SO}_4 + \text{C} \rightarrow \text{CO}_2 + \text{SO}_2 + \text{H}_2\text{O}$
- $\text{MnO}_2 + \text{HCl} \rightarrow \text{Cl}_2 + \text{MnCl}_2 + \text{H}_2\text{O}$
- $\text{KMnO}_4 + \text{NaI} + \text{H}_2\text{SO}_4 \rightarrow \text{I}_2 + \text{MnSO}_4 + \text{Na}_2\text{SO}_4 + \text{K}_2\text{SO}_4 + \text{H}_2\text{O}$
- $\text{K}_2\text{Cr}_2\text{O}_7 + \text{KBr} + \text{H}_2\text{SO}_4 \rightarrow \text{Br}_2 + \text{K}_2\text{SO}_4 + \text{Cr}_2(\text{SO}_4)_3 + \text{H}_2\text{O}$
- $\text{KMnO}_4 + \text{H}_2\text{C}_2\text{O}_4 + \text{H}_2\text{SO}_4 \rightarrow \text{MnSO}_4 + \text{K}_2\text{SO}_4 + \text{CO}_2 + \text{H}_2\text{O}$
- $\text{K}_2\text{Cr}_2\text{O}_7 + \text{SnSO}_4 + \text{H}_2\text{SO}_4 \rightarrow \text{K}_2\text{SO}_4 + \text{Cr}_2(\text{SO}_4)_3 + \text{Sn}(\text{SO}_4)_2 + \text{H}_2\text{O}$
- $\text{S}_2\text{O}_3^{2-} + \text{Cr}_2\text{O}_7^{2-} + \text{H}^+ \rightarrow \text{SO}_4^{2-} + \text{Cr}^{3+} + \text{H}_2\text{O}$
- $\text{N}_2\text{H}_4 + \text{Cu}(\text{OH})_2 \rightarrow \text{N}_2 + \text{Cu} + \text{H}_2\text{O}$ (Basic)